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284 Chapter 9 Chemical Reactions Figure 9.3 Science, like all other disciplines, has a specialized language that allows specific information to be communicated in a uniform manner. This reaction between aluminum and bromine can be described by a word equation, a skeleton equation, or a balanced chemical equation.

Chapter 9: Chemical Reactions

e. molar mass of $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$ $272 \text{ g Fe}(\text{NO}_3)_3$ $1 \text{ mol } 241.9 \text{ g} = 1.12 \text{ mol Fe}(\text{NO}_3)_3$ f. molar mass of $\text{FeSO}_4 = 151.9 \text{ g}$ 12.7 mg FeSO_4 $1 \text{ g } 1000 \text{ mg } 1 \text{ mol } 151.9 \text{ g} = 8.36 \times 10^{-5} \text{ mol FeSO}_4$ g. molar mass of $\text{LiOH} = 23.95 \text{ g}$ $6.91 \text{ } 103 \text{ g LiOH}$ $1 \text{ mol } 23.95 \text{ g} = 288.5 = 289 \text{ mol LiOH}$ 14. a. molar mass of $\text{CaCO}_3 = 100.1 \text{ g}$ $2.21 \times 10^4 \text{ mol CaCO}_3$ $100.1 \text{ g } 1 \text{ mol} = 0 \dots$

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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